



GCE MARKING SCHEME

**CHEMISTRY
AS/Advanced**

SUMMER 2012

CH2

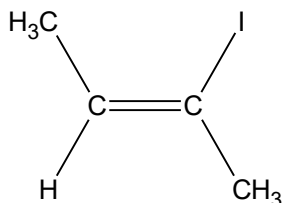
SECTION A

Q.1 (a) $C_{19}H_{40}$ [1]

(b) $C_{19}H_{40} \rightarrow C_8H_{18} + C_{11}H_{22}$ - allow ecf [1]

Q.2 2-chlorobutane [1]

Q.3 [1]



Q.4 any number in range 1 to 6 [1]

Q.5 (a) maximum mass = 44-45 (g) [1]

(b) (less solute would form as a solid) because more will remain in the solution [1]

Q.6 (a) iodine force is Van der Waals/ induced dipole-induced dipole (1)

diamond force is covalent bond/ description of attractive forces in a covalent bond (1) [2]

(b) diamond would have a higher sublimation temperature because it has stronger forces/ forces are harder to break [1]

Section A Total [10]

SECTION B

- Q.7 (a) (i) one σ bond/ description of σ bond/ diagram to show overlap of s orbitals (1)
- one π bond/ description of π bond/ diagram to show sideways overlap of p orbitals (1) [2]
- (ii) joining of **many/lots of** (small) units or many alkenes / molecules to make a **large/long** unit/ molecule [1]
- (iii)
- $$\left(\begin{array}{cc} \text{H} & \text{CH}_3 \\ | & | \\ \text{---C} & \text{---C---} \\ | & | \\ \text{H} & \text{CO}_2\text{CH}_3 \end{array} \right)_n$$
- [1]
- (iv) $\text{C}_4\text{H}_5\text{Cl}$ [1]
- (b) (i) BF_3 is planar triangular/ trigonal planar (1)
- NH_3 is pyramidal/ trigonal pyramid (1) [2]
- (ii) BF_3 has 3 bond pairs (1)
- NH_3 has 3 bond pairs and 1 lone pair (1) [2]
- QWC the information is organised clearly and coherently, using specialist vocabulary where appropriate* [1]
- (c) (i) co-ordinate/ dative covalent/ dative
- no credit for 'covalent' [1]
- (ii) $109\frac{1}{2}^\circ$ (accept any in range 109° - 110°) [1]
- (iii) 4 bond pairs/ bonds (around B)
- no credit for 'tetrahedral' [1]
- Total [13]

Q.8 (a) (i) % H = 14.3 (1)

$$\text{C} : \text{H} = \frac{85.7}{12.0} : \frac{14.3}{1.01} = 7.14 : 14.16 \text{ (1)}$$

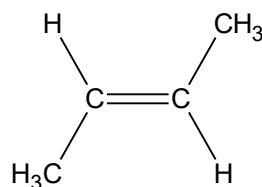
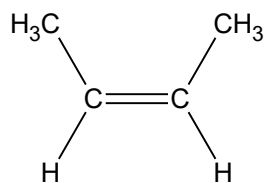
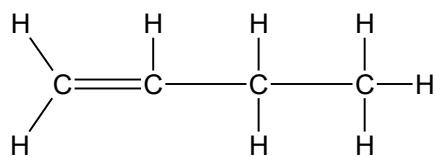
empirical formula = CH_2 (1) [3]

(ii) $M_r = 42$ / largest fragment has mass 42 (1)

($\text{CH}_2 = 14$) therefore molecular formula = C_3H_6 (1) [2]

(iii) CH_3 is present [1]

(b) 1 mark for each [3]



Total [9]

- Q.9 (a) apparatus in which reaction can occur, e.g. flask/ test tube, and delivery/ rubber tube (1)
- apparatus in which to measure volume of gas, e.g. over water with measuring cylinder/ gas syringe (1) [2]
- (b) (i) fewer **moles** of barium used / barium has a higher A_r [1]
- (ii) reaction faster/ more vigorous/ less cloudy solution formed with barium (1)
- because ionisation energy of barium is less/ electrons lost more easily from barium/ barium is lower in the group/
barium hydroxide is more soluble (1) [2]
- (c) flame test (1) brick red for calcium **and** (apple) green for barium (1)
- OR
- add sulfuric acid/ sodium sulfate solution/ potassium sulfate solution (1)
- white precipitate with Ba^{2+} , less precipitate/ no precipitate with Ca^{2+} (1) [2]
- (d) electrons correct – oxide ion clearly shows that 2 electrons originated from calcium atom (1)
- charges correct (1) [2]
- (e) (i) add sulfuric acid/ sodium sulfate solution/ potassium sulfate solution (1)
- filter (1)
- $Ba^{2+} + SO_4^{2-} \rightarrow BaSO_4$ (1)
- state symbols ignored [3]
- (ii) moles Ba = $2/137$ (1)
- mass $BaSO_4 = \frac{2 \times 233.1}{137} = 3.4$ (g) (1) [2]
- Total [14]

Q.10 (a) both contain metallic bonds/ positive ions and delocalised electrons labelled on diagram (1)

those in magnesium are stronger/ harder to break/ need more energy to break (1)

because **2** electrons are involved in delocalisation/ attraction to the positive ions (1) [3]

(b) reaction is hydrolysis of halogenoalkane/ nucleophilic substitution of halogenoalkane (1)



(white precipitate is) silver chloride and (cream precipitate is) silver bromide (1)



- state symbols ignored [4]

QWC selection of form and style of writing appropriate to purpose and to complexity of subject matter [1]

(c) caesium ions are bigger than sodium ions – accept ‘atoms’ (1)

co-ordination number 6 : 6 for sodium and 8 : 8 for caesium (1)

both cubic (1) [3]

(d) reaction is electrophilic addition (1)

two possible products are 1-bromopropane and 2-bromopropane (1)

more 2-bromopropane formed (1)

because of greater stability of intermediate positive ion/ 2° carbocation (1)

[4]

QWC legibility of text; accuracy of spelling, grammar and punctuation, clarity of meaning [1]

Total [16]

Q.11 (a) diagram completed with at least 1 water molecule and indication of interaction between O on one molecule and H on the other (1)

interaction between δ^+ on H and lone pair on O (1)

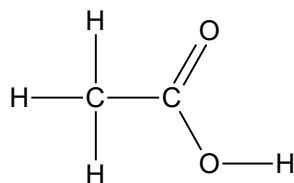
interaction labelled hydrogen bond (1) [3]

(b) (i) reduction/ redox – accept 'oxidation' [1]

(ii) I OH [1]

II OH is also present in water [1]

(c) (i) [1]



(ii) peak at 1650-1750 (1)

due to C=O (1) [2]

Total [9]

- Q.12 (a) incomplete p sub-shell/ outer electron configuration s^2p^5 / outer electrons in p subshell/ outer electrons in p orbitals/ valence electrons in p subshell/ valence electrons in p orbital [1]
- (b) (i) gaining one electron completes shell/ gives p^6 / takes an electron from another species/gains an electron
- do not accept 'attracts an electron' [1]
- (ii) **fluorine** because it is the smallest/ has the greatest electron affinity/ has the least shielding/ has the greatest effective nuclear charge/ oxidising power decreases as the group is descended [1]
- (c) oxidation state is (+)5/ V
- do not accept '5+' [1]
- (d) (i) $Cl_2 \rightarrow 2Cl^\bullet$ - ignore hf [1]
- (ii) $CH_4 + Cl^\bullet \rightarrow HCl + \bullet CH_3$ (1)
 $\bullet CH_3 + Cl_2 \rightarrow CH_3Cl + Cl^\bullet$ (1) [2]
- (e) products: $\bullet CFH_2$ and Cl^\bullet (1)
C-Cl bond is the **weakest/ most** easily broken (1) [2]

Total [9]

Section B Total [70]