

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the May/June 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/35

Paper 31 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.

Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	35

Question	Sections	Indicative material	Mark	
1 (a)	PDO layout	I Volume given for rough titre and accurate titre details tabulated <i>Minimum of 2×2 “boxes”</i>	1	
	MMO collection	II Follows instructions – dilutes 44.50–45.50 cm ³ FA 2 and records unambiguous initial and final burette readings and volume of FA 2 diluted and volume of FA 3 added for each titration. <i>Headings should match readings.</i> <i>Do not award this mark if:</i> <i>50.(00) is used as an initial burette reading;</i> <i>more than one final burette reading is 50.(00);</i> <i>any burette reading is greater than 50.(00)</i>	1	
	MMO decisions	III All accurate burette readings (initial and final) recorded to nearest 0.05 cm ³ including dilution table <i>Assess this mark on burette readings only, ignore volume of FA 3 added.</i>	1	
	PDO recording	IV has two titres within 0.10 cm ³ <i>Do not award this mark if having performed two titres within 0.1 cm³ a further titration is performed which is more than 0.10 cm³ from the closer of the initial two titres, unless a fourth titration, within 0.1 cm³ of any other has also been carried out.</i>	1	
	<p>Examiner to check and correct (if necessary) subtractions in the titre table. Examiner then selects the “best” titre using the hierarchy: two identical; titres within 0.05 cm³, titres within 0.10 cm³, etc., (ignore rough titre) For candidates and Supervisor scale titre for 45.00 cm³ FA 2 diluted. Calculate titre × $\frac{45.00}{\text{volume of FA 2 diluted}}$ to 2 dp Calculate difference in Supervisor and candidate scaled values and award “quality” marks as below.</p>			
	MMO quality	Award V , VI and VII for a difference from Supervisor, $\delta \leq 0.30 \text{ cm}^3$	1	
		Award V and VI for $0.30 < \delta \leq 0.60 \text{ cm}^3$	1	
		Award V only for $0.60 < \delta \leq 1.00 \text{ cm}^3$ <i>If “best” titres are 0.60 cm^3 apart cancel one of the Q marks</i>	1	[7]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	35

(b)	ACE interpretation	<p>Calculates the mean, correct to 2 decimal places from any accurate titres within 0.20 cm^3. <i>The third decimal place may be rounded to the nearest 0.05 cm^3.</i> <i>A mean of exactly $.x25$ or $.x75$ is allowed but the candidate may round up or down to the nearest 0.05 cm^3.</i> <i>If ALL burette readings are given to 1 decimal place then the mean can be given to 1 decimal place if numerically correct without rounding.</i> <i>Mean of 24.3 and 24.4 = 24.35 (✓)</i> <i>Mean of 24.3 and 24.4 = 24.4 (✗)</i></p> <p>Titres to be used in calculating the mean must be clearly shown – in an expression or ticked in the titration table.</p>	1	[1]
(c)	ACE interpretation	<p>I Expression correct in step (i) $\text{volume diluted} /_{250} \times 1.00$</p>	1	
		<p>II Correctly uses $\text{titre from (b)} /_{1000} \times \text{ans to (i) in (ii)}$ and $\frac{1}{2} \times \text{ans to (ii) in (iii)}$</p>	1	
		<p>III $\text{ans to (iii)} \times \frac{1000}{25} \times 201.2$ in (iv)</p>	1	
		<p>IV Uses $\frac{(38.10 - \text{ans to (iv)})}{38.10} \times 100$ in (v)</p>	1	
	PDO display	<p>V Working shown in all steps attempted and a minimum of 3 steps. (use of 2 in (iii), missing $\times 40$ or M_r in (iv) gains the mark) <i>(Working should be a step in the right direction)</i></p>	1	
		<p>VI 3 to 4 significant figures shown in final answer to all steps attempted – minimum of 3 steps</p>	1	[6]
(d)	ACE interpretation	<p>Correctly evaluates: $\frac{0.06}{25} \times 100$ or 0.24% and $\frac{0.10}{\text{titre in (b)}} \times 100$ <i>Answers must be given to at least 2 significant figures and correctly rounded for the significant figures shown.</i></p>	1	[1]
			[Total: 15]	

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	35

2 (a)	PDO layout	I All data presented clearly in all three sections. (6,6,7)	1	
	PDO recording	II Has correct headings and units on page 7.	1	
		III All thermometer readings recorded to nearest 0.5 °C in each of the experiments	1	
		IV Each pair of balance readings consistent and to at least 1 decimal place	1	[4]
(b)	Examiner to calculate (corrected) $\Delta T_1/m_1$ and $\Delta T_2/m_2$ for Supervisor and candidate. Compare candidate value with the same value from the Supervisor report. Award Q marks on the closer value.			
	MMO	Award I and II for $\delta < 0.10 \text{ } ^\circ\text{Cg}^{-1}$	1	
	quality	Award I only for $0.10 < \delta < 0.30 \text{ } ^\circ\text{Cg}^{-1}$	1	[2]
(c)	MMO collection	I Follows instructions – weighs between 8.5 and 9.5 g of FA 6 (mass bottle with FA 6 – mass bottle)	1	
	PDO layout	II Check Δm and ΔT are correct in (c)	1	[2]
(d)	ACE interpretation	Examiner to check there is no obvious error in the evaluation of the expression, then award one mark for a mass of sodium carbonate between 2.5 and 3.5 g.	1	[1]
(e)	ACE improvements	Give one mark for: suggesting weighing, heating and weighing again, or weighing, heating and measuring gas volume or giving an outline for a titration method using 2 indicators.	1	[1]
			[Total: 10]	

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	35

FA 8 is NaCl(aq); FA 9 is NaNO ₂ (aq); FA 10 is NaBr(aq); FA 11 is CuSO ₄ (aq); FA 12 is MgSO ₄ (aq)					
3	(a)	MMO decisions	Selects any named acid	1	
		MMO collection	Records brown gas with FA 9 and no reaction with FA 8 and FA 10	1	[2]
	(b)	MMO decisions	<p>I Selects: (<i>correct full name or formula</i>) silver nitrate as first reagent, aqueous ammonia as second reagent, aqueous ammonia added to tube with Ag⁺, 1st box ticked (<i>do not allow if Pb²⁺ used as 2nd reagent</i>)</p> <p>or</p> <p>lead nitrate as first reagent, silver nitrate as second reagent, Ag⁺(aq) added to fresh sample, 2nd box ticked</p>	1	
		MMO collection	<p>II <u>If Ag⁺ used as 1st reagent</u> Give one mark for white ppt with FA 8 and cream ppt with FA 10 <u>If Pb²⁺ used as 1st reagent</u> Give one mark for white ppt with FA 8 and FA 10 <i>If FA 9 not previously identified then no change/no reaction/no ppt (ignore any yellow colouration of solution with Pb²⁺)</i></p>	1	
			<p>III <u>If Ag⁺ used as 1st reagent (with NH₃ as 2nd)</u> Give one mark if white ppt with FA 8 is soluble in aqueous ammonia and cream ppt with FA 10 is insoluble or partially soluble in aqueous ammonia <u>If Ag⁺ used as 1st reagent (with Pb²⁺ as 2nd)</u> Allow observations marks <u>If Pb²⁺ used as 1st reagent (with Ag⁺ as 2nd)</u> Give one mark for white ppt with FA 8 and Ag⁺ and cream ppt with FA 10 and Ag⁺. <i>Ignore observations for FA 9.</i></p>	1	[3]
	(c)	ACE conclusion	Mark consequentially on observations; Give one mark for appropriate anions identified for FA 8 , FA 9 and FA 10 . (Allow from off-white or cream ppt for Br + Ag ⁺)	1	[1]

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – May/June 2011	9701	35

(d)	PDO recording	I Observations in a single table. All additions of NaOH(aq) and NH ₃ (aq) shown to excess where there is an initial ppt	1	
	MMO collection	II All observations correct for FA 11 (Blue ppt in each, blue ppt insoluble in excess NaOH, soluble in excess NH ₃ or forming/turning to a deep/dark blue solution)	1	
		III All observations correct for FA 12 (White ppt insoluble in each)	1	[3]
(e)	ACE conclusion	I Mark consequentially to observations. Expected conclusion is Cu ²⁺ in FA 11 and Mg ²⁺ in FA 12 Allow Ca ²⁺ from white ppt insoluble in excess NaOH and no ppt with NH ₃ .	1	
		II Gives appropriate evidence for each ion in the conclusion. Minimum evidence required for the expected ions: Cu²⁺ Records a blue ppt with either of the reagents or deep blue solution with excess NH ₃ . Mg²⁺ White ppt insoluble in excess NH ₃ (or in each of the reagents)	1	[2]
(f)	MMO collection	I Blue, black, purple colour observed on adding starch in (ii)	1	
		II The brown (solution) or (brown) solution formed in (i) is decolourised/colour fades/paler or brown (solution) in (i) and white, off-white or light brown ppt recorded.	1	
	ACE conclusion	Award III and IV for two correct pairs	1	
		Award III only for one correct pair Expected results (i) I is oxidised, Cu ²⁺ is reduced (ii) S ₂ O ₃ ²⁻ is oxidised, I ₂ is reduced Mark horizontally or vertically.	1	[4]
			[Total: 15]	