

Cambridge
International
AS & A Level

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Cambridge International Advanced Subsidiary and Advanced Level

CHEMISTRY

9701/21

Paper 2 AS Structured Questions

October/November 2016

MARK SCHEME

Maximum Mark: 60

Published

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This document consists of **8** printed pages.

Page 2	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	
1(a)	6×10^{-3} (mol)	1	1
1(b)	$\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$	1	1
1(c)	6×10^{-3} (mol)	1	1
1(d)	4×10^{-3} (mol)	1	1
1(e)	4×10^{-3} (mol)	1	1
1(f)	1×10^{-3} (mol)	1	1
1(g)	170	1	1
1(h)	28(.0) Si/silicon	1 1	2
	Total:	9	

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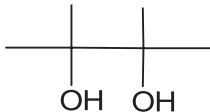
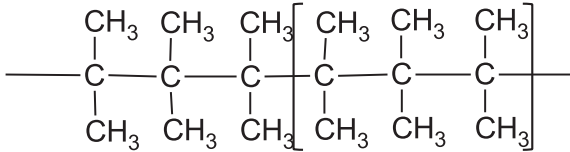
Question	Answer	Marks
2(c)(ii)	$\Delta H_f + (-2021.0) = 3(-393.5) + 4(-285.8)$ $\Delta H_f = -302.7 \text{ (kJ mol}^{-1}\text{)}$	1 2 1
	Total:	13

Question	Answer	Marks
3(a)(i)	(Atoms/ ions become larger as) the number of (electron) shells increases (down the group) Increased distance of (outer) electrons (from the nucleus) OR Increased shielding results in weaker (nuclear) attraction/pull	1 2 1
3(a)(ii)	top line / dotted line is atomic radii / bottom line / line with crosses is ionic radii (as atoms bigger than ions) Atom has one more shell (than corresponding ion) (ora) OR Atom loses two electrons / outer (shell) electrons / valency electrons (ora) OR Atom loses electrons and so (nuclear) attraction is stronger OR Nuclear charge in ion is greater than the electron(ic) charge (ora) OR Effective nuclear charge in ion is greater (ora)	1 2 1
3(b)(i)	Nitrate / Nitrate(V) / NO_3^-	1 1
3(b)(ii)	Ba / barium OR Sr / Strontium $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4$ OR $\text{Sr}^{2+} + \text{SO}_4^{2-} \rightarrow \text{SrSO}_4$	1 1

Page 5	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	
3(b)(iii)	$\text{Ba}(\text{NO}_3)_2$ OR $\text{Sr}(\text{NO}_3)_2$ $2\text{Ba}(\text{NO}_3)_2 \rightarrow 2\text{BaO} + 4\text{NO}_2 + \text{O}_2$ OR $2\text{Sr}(\text{NO}_3)_2 \rightarrow 2\text{SrO} + 4\text{NO}_2 + \text{O}_2$	1 1	2
3(c)(i)	$\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ OR $\text{Ca}(\text{OH})_2 + 2\text{H}^+ \rightarrow \text{Ca}^{2+} + 2\text{H}_2\text{O}$ $2\text{H}^+ + \text{CO}_3^{2-} \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ OR $\text{CaCO}_3 + 2\text{H}^+ \rightarrow \text{Ca}^{2+} + \text{CO}_2 + \text{H}_2\text{O}$ OR $\text{H}^+ + \text{CO}_3^{2-} \rightarrow \text{HCO}_3^-$ OR $\text{CaCO}_3 + \text{H}^+ \rightarrow \text{Ca}^{2+} + \text{HCO}_3^-$	1 1	2
3(c)(ii)	Calcium carbonate is insoluble /less soluble (ora) Calcium carbonate is less likely to be / won't get washed away (ora) OR Calcium carbonate lasts longer (ora) OR Calcium carbonate is less reactive / reacts more slowly (ora)	1 1	2
3(d)	$\text{Mg}(\text{OH})_2$ MgO	1 1	2
	Total:		14

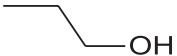
Page 6	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	
4(a)(i)	4-methylhex-2-ene	1	1
4(a)(ii)	(Molecules with the) same structural formula (and same molecular formula) with different arrangement of atoms/groups (in space)	1	1
4(a)(iii)	4 double-bond / alkene (2) different groups on each double-bonded carbon (one) chiral carbon (centre) / (one) carbon atom has 4 different groups attached / is asymmetric / is chiral	1 1 1 1	4
4(b)(i)	2,3-dimethylbut-2-ene	1	1
4(b)(ii)		1	1
4(b)(iii)	Propanone	1	1
4(b)(iv)		1	1
4(c)(i)	(2-)methylprop(-1-)ene	1	1

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	
4(c)(ii)		4	4
4(c)(iii)	(tertiary carbocat)ion / (tertiary) intermediate is / C⁺ with least number of hydrogen atoms bonded to it is more stable (than primary) due to (positive) inductive effect of three/more methyl groups (cf one) / three / more electron releasing methyl groups three / more electron donating methyl groups reducing charge (density) on C⁺	1 1 1	3
	Total:		18

Page 8	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks	
5(a)		1	1
5(b)	$\text{H}^+ / \text{Cr}_2\text{O}_7^{2-}$ (heat under) reflux	1 1	2
5(c)	$\text{H}^+ / \text{Cr}_2\text{O}_7^{2-}$ (heat and) distil	1 1	2
5(d)	(1-)propyl propanoate	1	1
	Total:		6